

Chapter 6 Review Chemical Bonding

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6.2 Covalent Bonding and Molecular Compounds Ionic and Covalent Bonds Made Easy

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CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

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1. _____ A chemical bond between atoms results from the attraction between the valence electrons and _____ of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2.

CHAPTER 6 REVIEW Chemical Bonding

CHAPTER 6 REVIEW Chemical Bonding SECTION 5 SHORT ANSWER Answer the following questions in the space provided. 1. Identify the major assumption of the VSEPR theory, which is used to predict the shape of atoms. Pairs of valence

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electrons repel one another. 2. In water, two hydrogen atoms are bonded to one oxygen atom. Why isn ' t water a linear molecule?

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Chemical bonding that results from the electrical attraction between positive ions and negative ions is called ionic bonding. In a purely ionic bond, the metal atom gives up its electron or electrons to the nonmetal atom. In covalent bonding, a bond forms from the sharing of electron pairs between two atoms. In a purely covalent bond, the

CHAPTER 6 Chemical Bonding

Section 6.2 – Covalent Bonding A covalent bond is a chemical bond in which two atoms share a pair of valence electrons. When two atoms share one pair of electrons, the bond is called a single bond. Covalent vs Ionic Bond

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Chapter 6 – Chemical Bonds

an attraction between atoms that allows the formation of chemical substances that contain two or more atoms. The bond is caused by the electrostatic force of attraction between opposite charges, either between electrons and nuclei, or as the result of a dipole attraction.

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CHAPTER 6 REVIEW Chemical Bonding SECTION 6-2 SHORT ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. 2. Name two elements that form compounds that are exceptions to the octet rule. 3.

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